

Problem-2

Two stable isotopes of lithium ${}_3\text{Li}^6$ and ${}_3\text{Li}^7$ have respective abundances of 7.5% and 92.5%. These isotopes have masses 6.01512 u and 7.01600 u, respectively. Find the atomic mass of lithium.

Solution:-

Atomic mass of lithium

$$= \left(\frac{7.5}{100} \times 6.01512 \right) + \left(\frac{92.5}{100} \times 7.01600 \right)$$

$$= 6.941 \text{ u}$$

-x-